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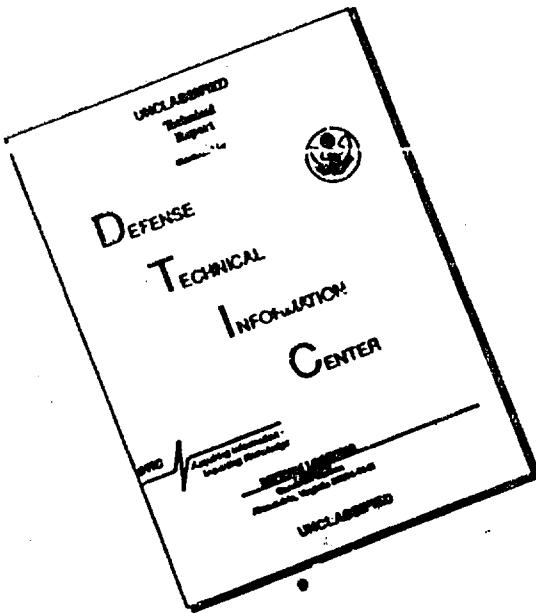
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P.T.P. 528

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Copy No. 80

MINISTRY OF SUPPLY

DIRECTORATE OF CHEMICAL DEFENCE RESEARCH AND DEVELOPMENT

CHEMICAL DEFENCE EXPERIMENTAL ESTABLISHMENT

FC

THE DETERMINATION OF FREE
BASE IN STABILISED GB.

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The Determination of Free Base in

Stabilised GB

by

E. Neale

SUMMARY

In view of possible loss of triethylamine from stabilised GB during filtration under reduced pressure the need has arisen for a method of determining the free base in the charging. A simple colorimetric method of estimating triethylamine has been developed and applied to the measurement of partial pressure of the base in low mole fractions over GB. Some information on the equilibria between triethylamine and the various acidic impurities in GB has been obtained by the same technique.

The deviations from ideality shown by the triethylamine/GB system are of practical interest in demonstrating that the volatility of the base will be greater than would be expected from the vapour pressure of the pure material.

The data on acid-base equilibria emphasise the desirability of keeping the acidity of GB low rather than relying too much on its 'neutralisation' by added base.

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The Determination of Free Base in
Stabilised GB

by

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Introduction

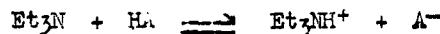
Although triethylamine is one of the most effective agents for 'neutralising' the acidic impurities in technical GB, and has for some time been accepted as the stabiliser for GB produced in this Country, there are certain difficulties associated with its use in large scale production;

- (a) Some of the impurities (e.g. hydrochloric acid) form salts with triethylamine which are insoluble in GB, thus necessitating a filtration system in the plant.
- (b) The volatility of triethylamine is such that any free base in the stabilised charging may be partly lost by evaporation when the material is filtered under reduced pressure. It is of course desirable to maintain a slight excess as further acidity may develop from hydrolysis or other decomposition in storage.

The primary object of the work reported below was to develop a simple method of determining excess triethylamine in GB which could be applied to the stabilised material after filtration. The investigation was subsequently extended to obtain information on acid-base equilibria in GB.

Choice of method

Reaction between the weaker acidic impurities and triethylamine is an equilibrium process:



Any direct chemical analysis of the charging is therefore precluded as removal of triethylamine will shift the equilibrium towards the left. As triethylamine is volatile however, a possible method of approach is to measure its partial pressure in the system since the salts formed may be regarded as involatile. Ideally the pressure should be measured by a physical method which does not disturb the equilibrium but this is difficult in practice; a good approximation may be obtained however by removing a vapour sample for analysis provided the amount of triethylamine in the vapour phase is small compared with that in the liquid.

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peak is at 5100 \AA . This afforded a simple colorimetric method of estimating microgram quantities. The presence of GB in the chloroform had no effect nor was the presence of small amounts of water sufficiently serious to justify any special measures to dry the chloroform. Details of procedure are given in Appendix 1 with a calibration curve as Fig. 2.

Materials

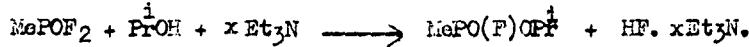
(a) Technical GB (SM.238) prepared at Sutton Oak by the 'di-di' process and stored in varnished steel for some years. Apparent acidity 490 p.p.m. This material was chosen to represent the worst possible acidities which might be encountered in stored commercial GB, though it is, of course, not representative of present production.

(b) Pure GB, from the technical grade - twice distilled giving an apparent acidity less than 1 p.p.m.

(c) Triethylamine: Hopkins and Williams product dried over barium oxide $N_D^{20} = 1.4001$. Diethylamine content not more than 0.1 per cent.

(d) Acidic Materials. Methyl phosphonofluoridic acid was prepared and purified (at the Ministry of Supply Establishment, Nancekuke) shortly before use. Methylphosphonic dichloride and isopropyl methylphosphonochloridate were distilled under reduced pressure as required. Methylphosphonic difluoride was distilled at atmospheric pressure over diethyl aniline. Methylphosphonic acid and its isopropyl half-ester had been prepared and purified some months before but were stored in a desiccator.

In view of the difficulty of adding known amounts of dry HF to GB, equimolar amounts of dry isopropanol and pure methyl phosphonic difluoride were added to GB containing an excess of triethylamine:



Analysis of a sample by the Poroxide method modified for elimination of $\text{MePOF}_2(5)$ showed that the conversion was quantitative within the limits of the analytical method.

Results

The partial pressure of triethylamine over pure GB at various low mole fractions is shown graphically in Fig. 3. Results obtained by the static and dynamic methods are compared with the theoretical curve for the system obeying Raoult's law. In calculating the latter, the vapour pressure data of Thompson and Linnett (6) are used.

It will be noted that there is a marked positive deviation from Raoult's Law. The fact that results by the dynamic method are somewhat higher than those by the static method needs some comment. A test with triethylamine alone in the dynamic apparatus gave a recovery equivalent to a saturated vapour concentration of 350 mg/l. The value calculated from the published vapour pressure data is 305 mg/l; the discrepancy may be due to adsorption of vapour on the surface of the glass. The static method is not open to this objection; here results are not likely to be high unless the air contains liquid droplets - this is improbable as similar results were obtained after passage through a glass wool filter. Furthermore the curve obtained by the static method may

be expressed well by an equation of the Duhem-Hargules type for a binary liquid system:

$$p_A' = p_A^0 x_A E^{\alpha x_B^2}$$

where x_A and x_B are the mole fractions of solute and solvent and p_A^0 is the vapour pressure of the pure solute. In the present case the constant α has a value of 1.18 which gives a measure of the extent to which the system deviates from the ideal.

Despite the evidence that results given by the dynamic method are high, this technique is preferable for routine determination of free base in GB for the following reasons:

- (i) reproducibility is very good and much superior to that given by the static method. If therefore the same apparatus is used for relating partial pressure to mole fraction of base in pure GB and for examination of unknown samples, the error due to possible adsorption of vapour on the glass will be eliminated.
- (ii) the procedure is much more convenient.

Measurements made on technical GB of high acidity are shown in Fig. 4. It will be noted that the apparent acidity value gives little indication of the amount of triethylamine which would react with the acid impurities.

Results of similar measurements on samples of pure GB to which known amounts of acidic materials had been added are summarised in Appendix 2 and graphed in Figs. 5 and 6. The amounts of triethylamine reacted with isopropyl hydrogen methylphosphonate ('half-ester') and with methyl phosphono-fluoridic acid ('fluoro acid') correspond fairly well with the formation of a 1:1 salt, the fluoro acid being the stronger as expected. The equilibrium constant (calculated on mole fractions and assuming the salt to be un-ionised) for the 'fluoro-acid' is approximately 9.3×10^2 compared with 2.8×10^2 for the 'half-ester'. Methyl phosphonic acid appears to be slightly stronger than the 'half-ester' and there is evidence of the second hydroxyl group co-ordinating slightly when triethylamine is present in excess. The curve for HF indicates that four moles of acid co-ordinate with three of base; this would not be inconsistent with the conductimetric results obtained by Wardrop and Bryant (7). The data for methyl phosphonic difluoride would correspond to the formation of $\text{Et}_3\text{N} \cdot 2 \text{MePOF}_2$ and that for the dichloride appears to be similar but results both with MePOCl_2 and with isopropyl methyl-phosphonochloride ('chloro GB') are suspect as trouble was experienced due to the precipitate clogging the capillary in the bubbler.

Conclusions

A simple method of estimating free triethylamine in GB has been developed and has been used to measure the partial pressure of triethylamine over GB in low mole fractions. Some data on acid base equilibria in GB have also been obtained.

The marked positive deviations from ideality shown by the triethylamine/GB system are of interest from the practical viewpoint in that the volatility of the base will be considerably greater than would be expected from the vapour pressure of the pure material.

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The implications of the data on acid-base equilibria can most readily be appreciated by considering the 'fluoro-acid' and 'half-ester' both of which are likely impurities which appear to form simple 1:1 salts with triethylamine. The approximate proportion of acid co-ordinated by varying amounts of amino are tabulated below:

<u>Et₃N added</u> (Mole per cent of acid present)	Per cent of acid present co-ordinated with Et ₃ N	
	Fluoro Acid	Half Ester
80	75	66
100	88	77
120	95	86

This emphasises the desirability of keeping the acidity of the GB low rather than relying too much on its 'neutralisation' by added base.

It would be of interest to extend this investigation to bases other than triethylamine.

Acknowledgements

Mrs. E. Letts gave assistance with the experimental work. Mr. P.J. Bryant purified the GB.

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Appendix 1 to P.T.P. 528

Colorimetric Estimation of Triethylamine

Stock Solution. One gram of iodine (A.R.) in 100 ml. chloroform (A.R.) may be kept indefinitely in the dark with a well stoppered bottle. For use, 10 ml. of stock solution is diluted to 100 ml. with chloroform.

Procedure. To a suitable aliquot of the sample (in chloroform) in a 25 ml. flask add 5 ml. of the 0.1 per cent iodine solution. Place in a bath at 25°C for 10 min., then leave for 5 min. to reach room temperature before making up to 25 ml. with chloroform. Measure on the Spekker using 601 Filters and a blank solution in the reference cell.

Interference. from volatile acidic compounds such as HF and MePOF_2 was overcome by first shaking the sample with 0.1N. NaOH saturated with chloroform. A test estimation with chloroform containing a known amount of Et_3N and MePOF_2 showed this procedure to give results which were correct to within five per cent.

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Appendix 2 to P.T.P. 528

Acid-Baso Equilibria in GB

1. Isopropyl hydrogen methyl phosphonate

Wt. of GB (gm)	Wt. of acid (gm)	Wt. of Et ₃ N (gm)	Mole Fractn. acid	Mole Fractn. Et ₃ N	Concn. of Et ₃ N in air (ng/l)	Equivt. Mole Fr. Et ₃ N in liquid
4.773	0.2710	0.1225	0.053	0.032	4.2	0.004
4.715	0.2652	0.1671	0.051	0.044	11.7	0.009
4.941	0.2618	0.3239	0.047	0.079	40.0	0.036
4.820	0.2643	0.4832	0.047	0.117	71.0	0.072

2. Methyl phosphonofluoridic acid

4.720	0.2441	0.1851	0.066	0.048	3.2	0.0025
4.830	0.2562	0.2667	0.066	0.066	8.8	0.008
4.618	0.2307	0.3015	0.061	0.078	23.2	0.020
4.705	0.2410	0.4132	0.061	0.102	46.0	0.0425

3. Methyl phosphonic acid

4.469	0.1059	0.1056	0.032	0.031	6.9	0.006
4.622	0.1095	0.2080	0.032	0.057	26.1	0.023
4.425	0.1048	0.3039	0.031	0.084	53.0	0.050

4. Hydrofluoric acid

5.197	0.039	0.1080	0.049	0.027	9.5	0.0085
5.100	0.042	0.2262	0.052	0.055	22.0	0.019
5.442	0.040	0.3307	0.046	0.075	44.0	0.040
5.096	0.043	0.5092	0.049	0.116	83.0	0.080

5. Methyl phosphonic difluoride

4.438	0.2098	0.1093	0.060	0.031	6.6	0.006
4.715	0.2240	0.2038	0.059	0.053	38.3	0.034
4.479	0.2695	0.3527	0.071	0.092	70.0	0.069
4.416	0.2224	0.3384	0.060	0.090	66.0	0.065

6. Isopropyl methylphosphonochloridato *

5.841	0.1446	0.1104	0.017	0.025	12.7	0.011
5.364	0.1366	0.2084	0.021	0.050	39.0	0.035

7. Methyl phosphonic dichloride *

4.645	0.2390	0.4150	0.046	0.106	82.0	0.080
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* Precipitate formed.

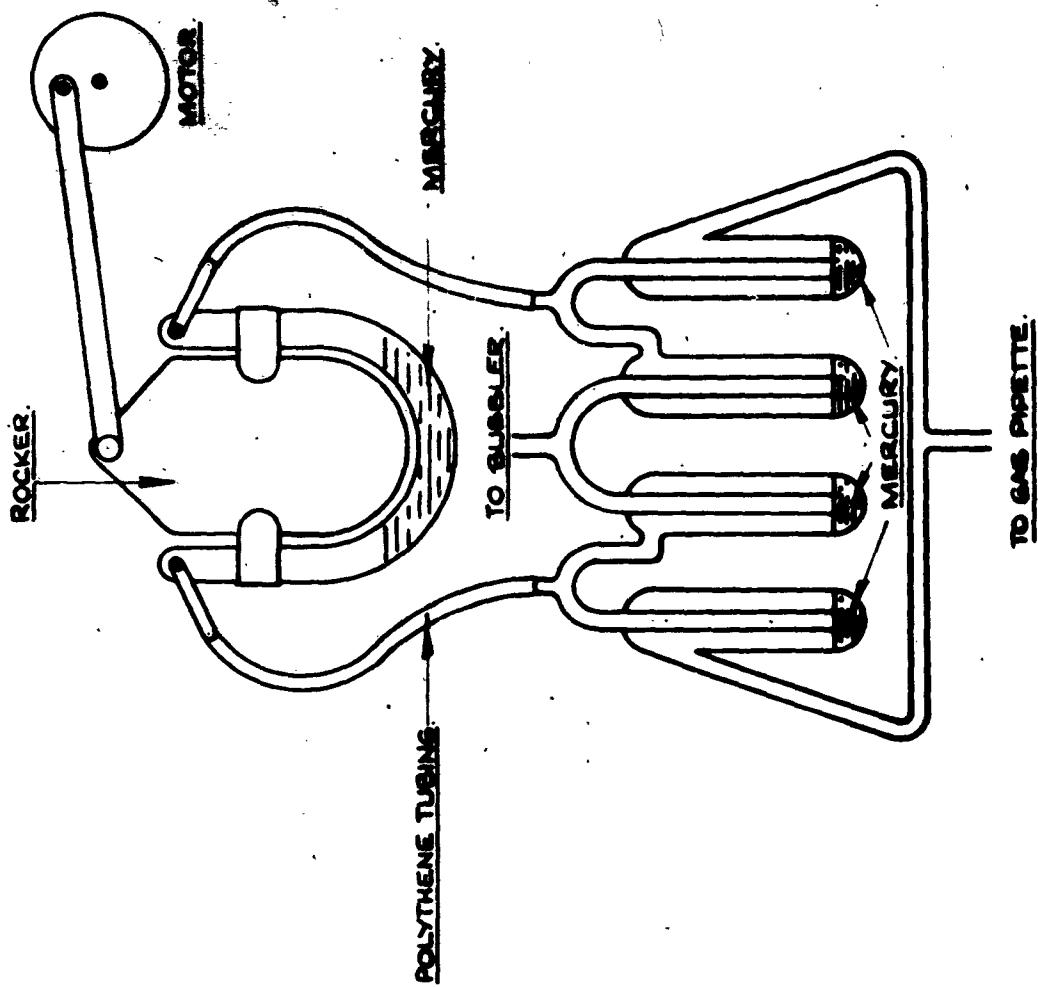


FIG. 1B. CIRCULATING PUMP.

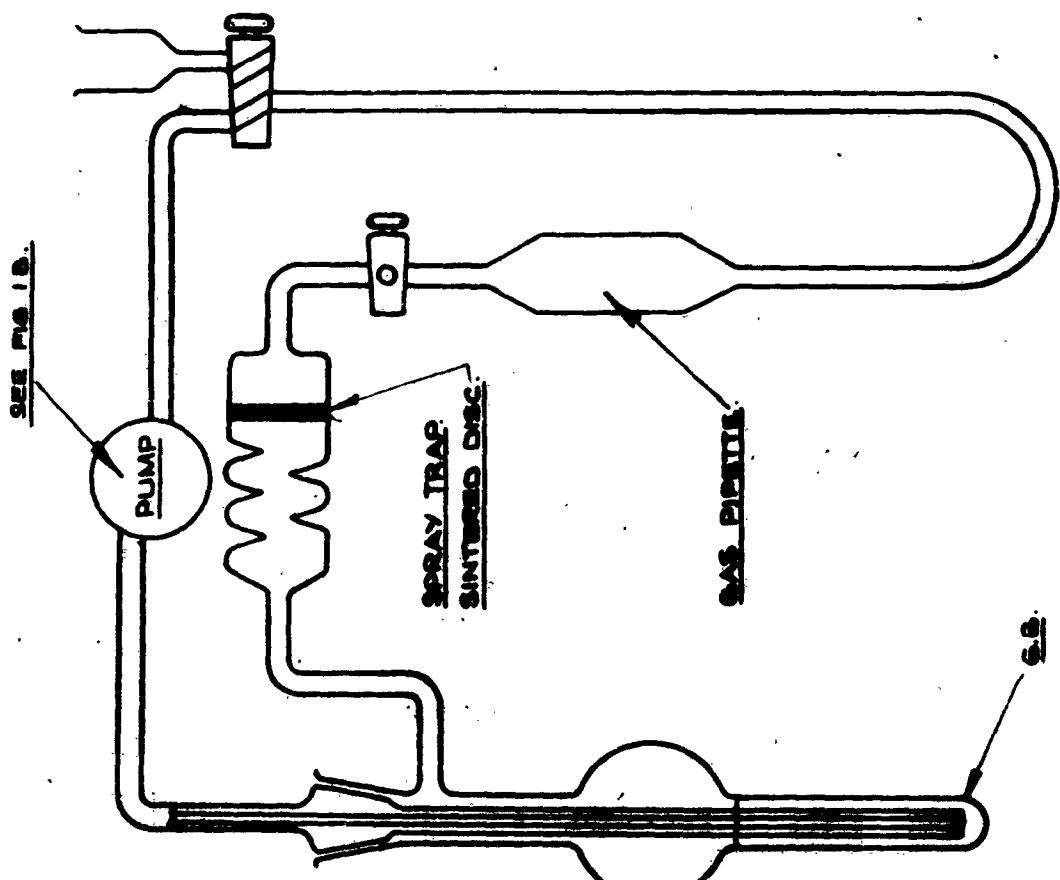


FIG. 1A. PARTIAL PRESSURE APPARATUS - DYNAMIC.

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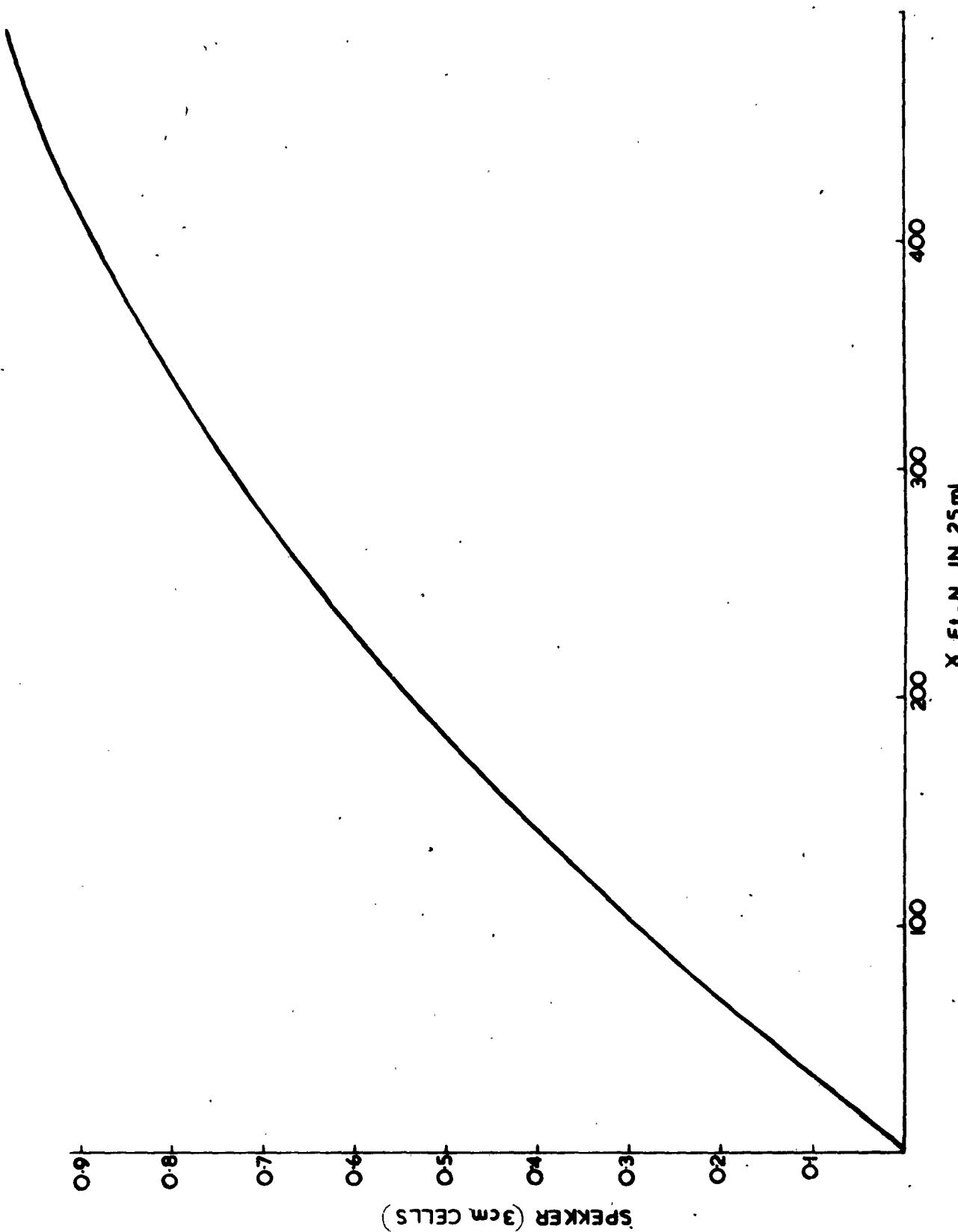
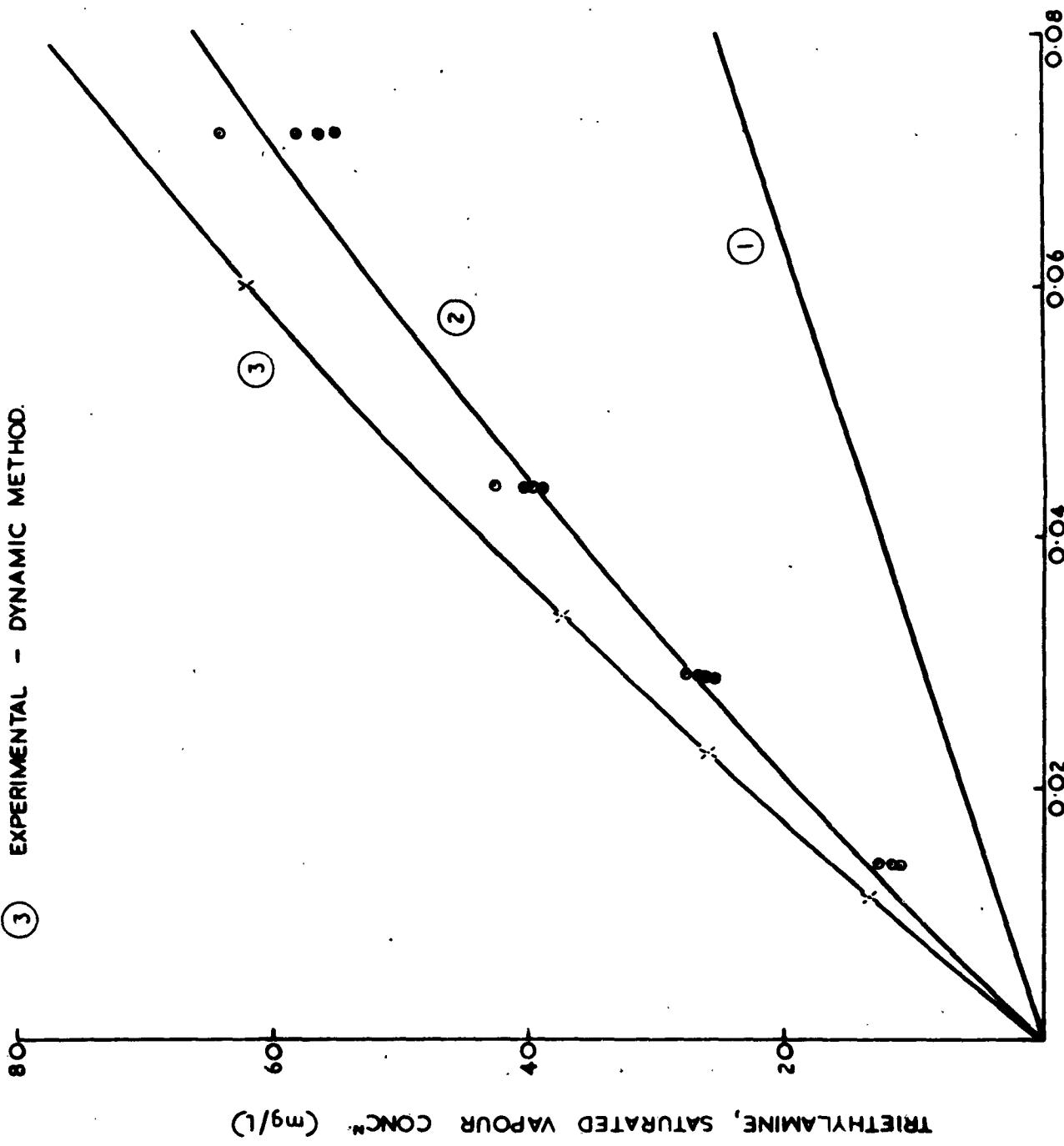


FIG.2. ESTIMATION OF $E_{1,N}$ WITH IODINE: CALIBRATION CURVE

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PARTIAL PRESSURE OF TRIETHYLAMINE OVER PURE GB (20°C)

- (1) THEORETICAL (RAOULTS LAW)
- (2) EXPERIMENTAL - STATIC METHOD.
- (3) EXPERIMENTAL - DYNAMIC METHOD.



MOLE FRACTION OF TRIETHYLAMINE.

FIG. 3

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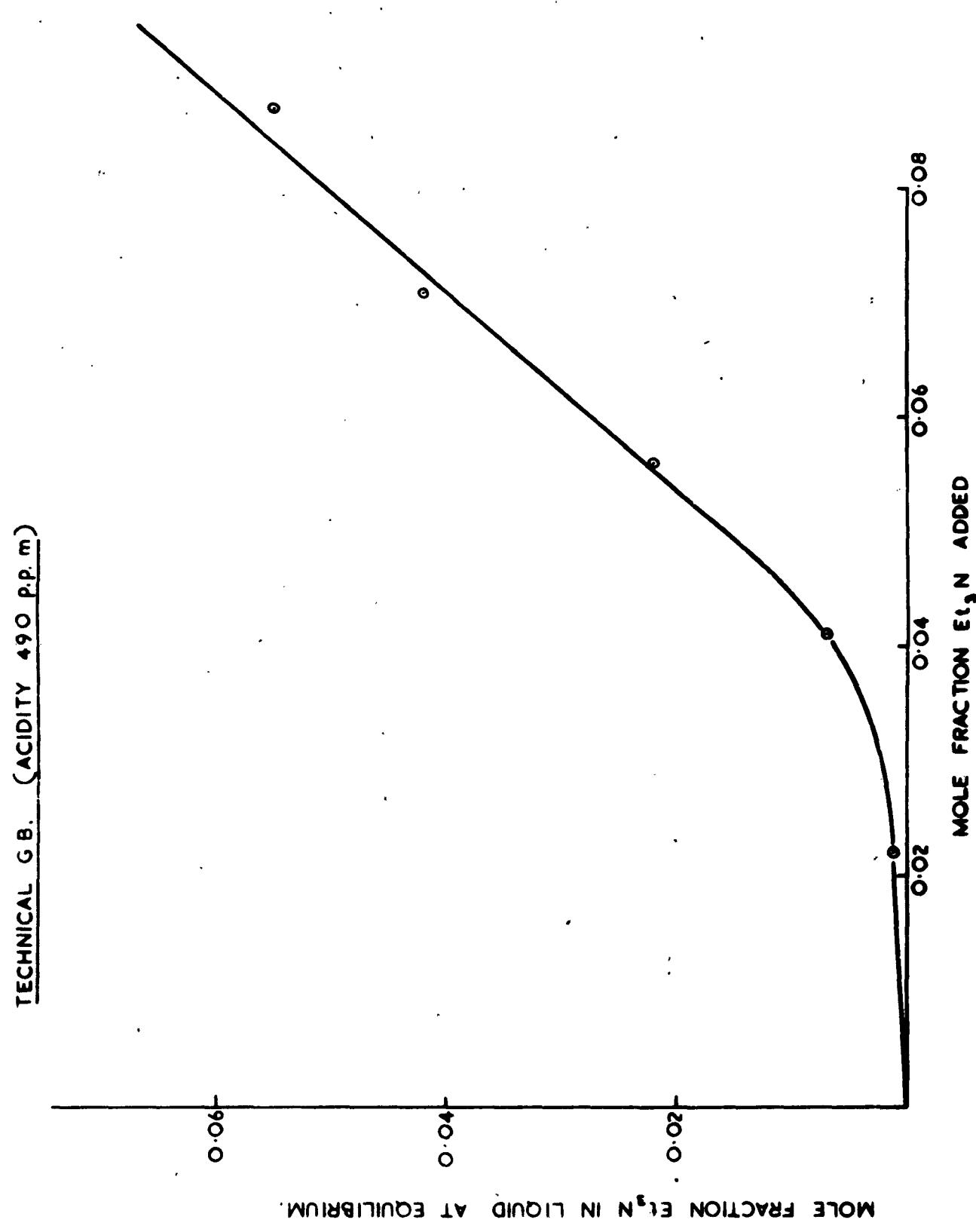


FIG. 4

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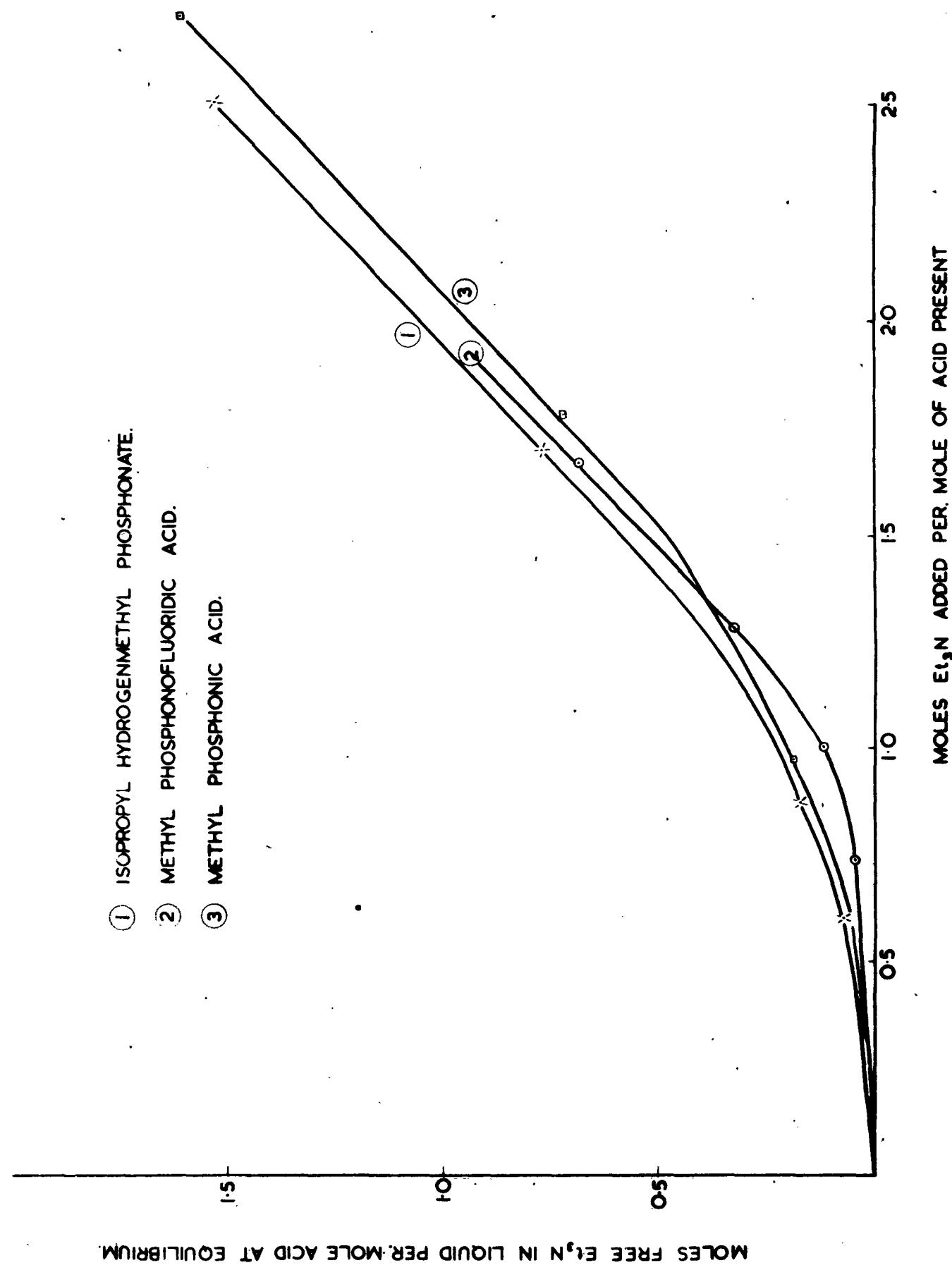


FIG. 5 ACID-BASE EQUILIBRIA IN G.B.

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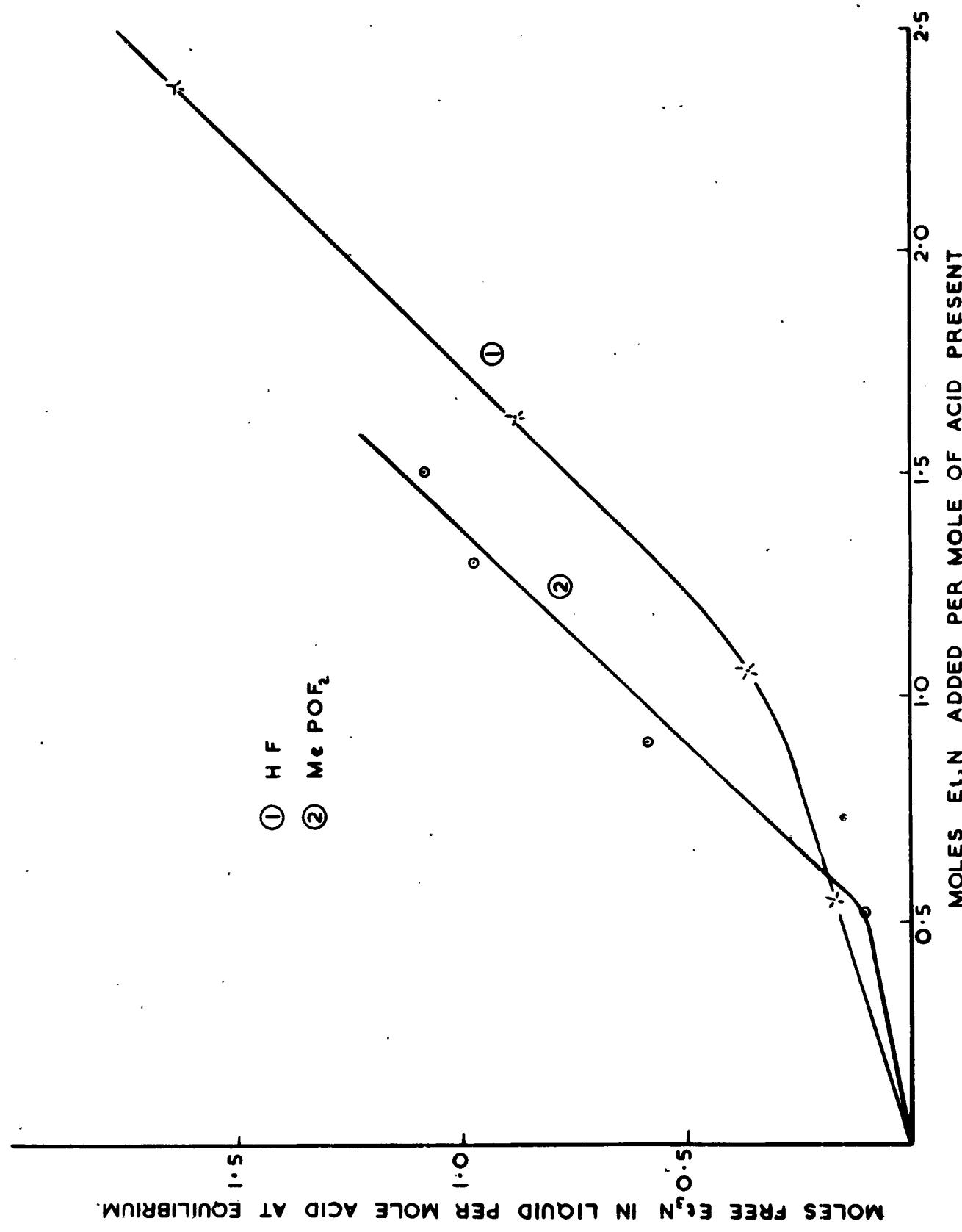


FIG. 6. ACID- BASE EQUILIBRIA IN G. B.

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